

Chemistry 1
Solution Practice

Name: _____
Period ____ Date _____

Partner #1 _____ Partner #2 _____ Partner #3 _____

Molarity = _____

Directions: Calculate the molarity of each of the following solution and show your work. Use correct units.

Set 1: Moles of Solute and Liters of Solution are given

1. 1.0mol of KCl in 0.750L of solution?
2. 0.50mol of MgCl_2 in 1.5L of solution
3. 0.060mol NaHCO_3 in 1.5L of solution

Set 2: Grams of Solute and Liters of Solution are given

4. 400 g of CuCl_2 in 4.00 L of solution
5. 12.6 grams HNO_3 in 1.0 L of solution
6. 12.2-grams of CaCl_2 , is dissolved in enough water to make 0.085 L of solution

Set 3: Mixed Mole and Mass Problems

7. 42.5 g NaCl in 375 mL of solution

8. 22.0 g CuCl_2 in 1000 mL of solution.

9. How many moles of sucrose are dissolved in 250 mL of solution if the solution concentration is 0.150 M?

Challenge: Which solution is more concentrated? Solution "A" contains 50.0 g of CaCO_3 in 500.0 mL of solution. Solution "B" contains 6.0 moles of H_2SO_4 in 4.0 L of solution. **SHOW WORK!**