

Calculate Average Atomic Mass

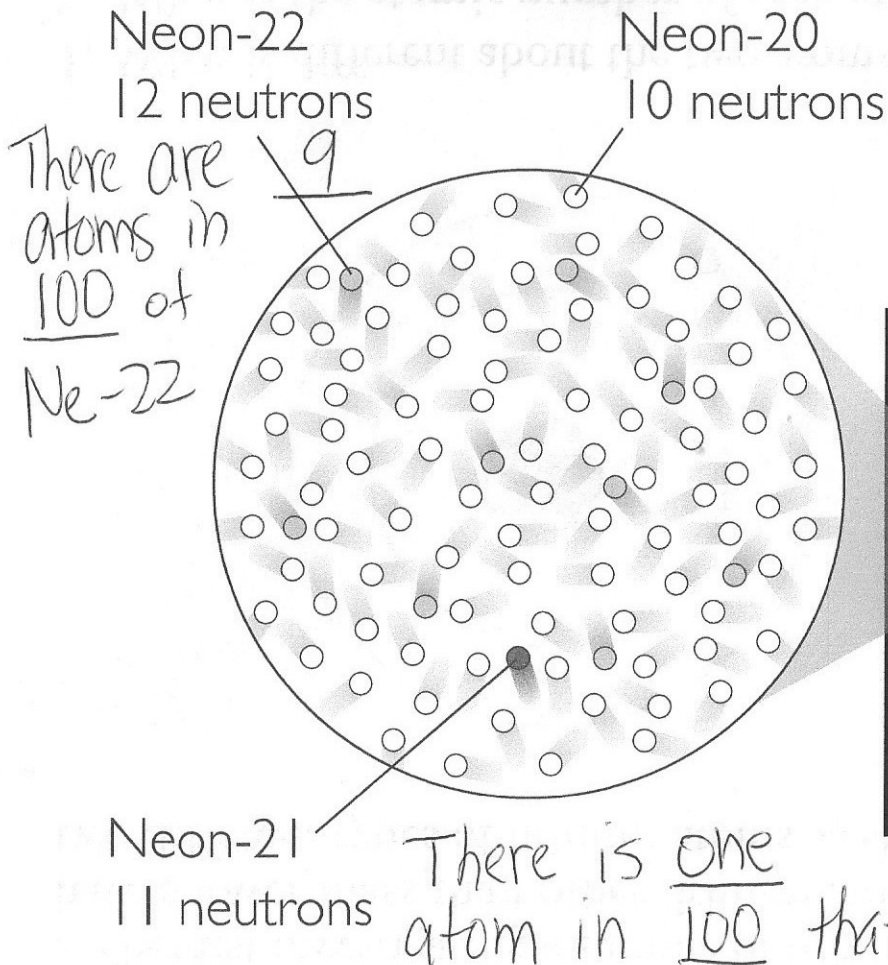
10/28/15

Abundance

Neon-21 $\frac{1}{100} = 0.01 \times 21 \text{ amu} = 0.21 \text{ amu}$
 Neon-22 $\frac{9}{100} = 0.09 \times 22 \text{ amu} = 1.98 \text{ amu}$
 Neon-20 $\frac{90}{100} = 0.90 \times 20 \text{ amu} = \underline{18.00 \text{ amu}}$

There are 90 atoms in 100 of
 of Ne-20 $[100 - 9 - 1 = 90]$.

Sample of 100 Neon Atoms



add these up

20.19 amu

Compare to
 P.T. Value = 20.18